

Adabistan-e-Soophia

Code: 1012

Test No.: 3

Paper: Chemistry

Name: _____

Class: IX Sec: _____

Syllabus: Ch. 3,8

Question Numbers	1	2	3			Total	Grade	%age
Maximum Marks	09	22	09			40		
Marks Obtained								

Remarks: _____

A	B	C	D	Write Correct option	A	B	C	D	Write Correct option	A	B	C	D	Write Correct option			
1	A	B	C	D		7	A	B	C	D		13	A	B	C	D	
2	A	B	C	D		8	A	B	C	D		14	A	B	C	D	
3	A	B	C	D		9	A	B	C	D		15	A	B	C	D	
4	A	B	C	D		10	A	B	C	D		16	A	B	C	D	
5	A	B	C	D		11	A	B	C	D		17	A	B	C	D	
6	A	B	C	D		12	A	B	C	D		18	A	B	C	D	

Note: Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more times result in zero mark in that question.

Q.1	Questions	(A)	(B)	(C)	(D)
1.	Melals lose this electrons easily because:	They are electro-negative	They have electron affinity	They are electro-positive	Good conductors of heat
2.	Sodium is extremely reactive metal, but it does not react with:	Hydrogen	Nitrogen	Sulphur	Phosphorus
3.	Metals can form ions carring charge:	Unipositive	Dipositive	Tri-positive	All of them
4.	The amount of energy given out when an electron is added to an atom is called:	Lattice energy	Ionization energy	Electro-negativity	Electron affinity
5.	Long form of periodic table is constructed on the basis of:	Mendeleev postulate	Atomic number	Atomic mass	Mass number
6.	Along the period, which one of the following decreases:	Atomic radius	Ionization energy	Electron affinity	Electro-negativity
7.	4 th and 5 th period of the long form of periodic table are called:	Short periods	Normal periods	Long periods	Very long periods
8.	Which one of the following halogen has lowest electronegativity?	Fluorine	Chlorine	Bromine	Iodine
9.	Which one of the following metal burns with brick red flame?	Sodium	Magnesium	Iron	Calcium

10.	Which one of the following is brittle?	Sodium	Aluminium	Selenium	Magnesium
11.	Which one of the following will not react with dilute <i>HCl</i> ?	Sodium	Potassium	Calcium	Carbon
12.	Non-metals are generally soft which one of the following is extremely hard?	Graphite	Phosphorus	Iodine	Diamond
13.	Which one of the following non-metal is lustrous?	Sulphur	Phosphorus	Iodine	Carbon
14.	Mark the incorrect statement about ionization energy.	It is measured in $KJ mol^{-1}$	It is absorption of energy	It decreases in a period	It decreases in a group
15.	Point out the incorrect statement about electron affinity.	It is measured in $KJ mol^{-1}$	It involves release of energy	It decreases in a period	It decreases in a group
16.	Transition elements are:	<i>s</i> block elements	<i>p</i> block elements	<i>d</i> block elements	All of these
17.	H. Moseley discovered a new property of the elements i.e atomic number in:	1913	1915	1900	1817
18.	Catalytic convertor is an alloy of:	Platinum, Palladium and Aluminium	Platinum, palladium and rhodium	Aluminium, iron and copper	Gold, silver and iron

(Section - I)

2. Write short answers to the following questions.**(11×2=22)**

- i. Define shielding effect. Describe its trend along the period.
- ii. Define atomic radius. Give example of atomic radius of Carbon.
- iii. Why are silver and gold least reactive?
- iv. What is relationship between electro-positivity and ionization energy?
- v. In bright sunlight, how Cl_2 and CH_4 react?
- vi. Define Electronegativity. Write its trend along group.
- vii. What is difference between ionization energy and electron affinity?
- viii. What is periodic law?
- ix. Write two physical properties of non-metals.
- x. Complete the following reactions:
 - a. $NaOH_{(dil)} + Cl_2 \rightarrow$
 - b. $Br_2 + H_2O \rightarrow$
- xi. What are transition elements?

(Section - II)

Note: Give detailed answers of the following questions.**(5+4=09)**

3. a) Write salient features of long form of periodic table. **(5)**
- b) Write down chemical properties of non-metals. **(4)**