

Adabistan-e-Soophia

Code: 2011

Test No.: 2

Paper: Chemistry

Name: _____

Class: X Sec: _____

Syllabus: Ch. 10

Question Numbers	1	2	3			Total	Grade	%age
Maximum Marks	09	22	09			40		
Marks Obtained								

Remarks: _____

Time Allowed: 15 mins

(Objective Type)

Max. Marks: 09

	A	B	C	D	Write Correct option		A	B	C	D	Write Correct option		A	B	C	D	Write Correct option
1	A	B	C	D		7	A	B	C	D		13	A	B	C	D	
2	A	B	C	D		8	A	B	C	D		14	A	B	C	D	
3	A	B	C	D		9	A	B	C	D		15	A	B	C	D	
4	A	B	C	D		10	A	B	C	D		16	A	B	C	D	
5	A	B	C	D		11	A	B	C	D		17	A	B	C	D	
6	A	B	C	D		12	A	B	C	D		18	A	B	C	D	

Note: Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more times result in zero mark in that question.

Q.1	Questions	(A)	(B)	(C)	(D)
1.	Dilute acids react with carbonates to produce the given products except:	Salt	water	Carbon dioxide	Hydrogen
2.	Lewis acid-base concept have the following characteristics except:	Formation of an adduct	Formation of a coordinate covalent bond	Donation and acceptance of an electron pair	Donation and acceptance of a proton
3.	The conjugate acid of HPO_4^{2-} is:	PO_4^{3-}	$H_2PO_4^{2-}$	$H_2PO_4^-$	H_3PO_4

4.	Which one of the following species is not amphoteric?	H_2O	NH_3	HCO_3^-	SO_4^{2-}
5.	Which one of the following is a Lewis base?	NH_3	BF_3	H^+	$AlCl_3$
6.	According to the Lewis concept, Acid in a substance which can:	Donate a proton	Donate a pair of electrons	Accept a proton	Accept a pair of electrons
7.	The product of Lewis acid-base reaction is called adduct. The bond between adduct specie is:	Ionic	Covalent	Metallic	Coordinate covalent
8.	A reaction between an acid and a base produces:	Salt and water	Salt and gas	Salt and an acid	Salt and a base
9.	A base is a substance which neutralizes an acid. Which of these substances is not a base?	Aqueous ammonia	Sodium chloride	Sodium carbonate	Calcium oxide
10.	If a liquid has a pH of 7 then it must:	Be a colorless and odourless liquid	Freezes at $0^\circ C$ and boils at $100^\circ C$	Be neutral	Be a solution containing water
11.	Acetic acid is used for:	Flavouring food	Making explosive	Etching designs	Cleaning metals
12.	When an acid reacts with metal, the gas evolves is:	CO_2	SO_2	CO	H_2
13.	Example of mineral acids:	Hydrochloric acid	Sulphuric acid	Nitric acid	All of these
14.	Sodium hydroxide is used for:	Manufacturing of bleaching powder	Manufacturing of soap	Removing grease and stains from clothes	Treatment of bee's stings
15.	The colour of methyl orange arrange in alkaline solution is:	Yellow	Pink	Red	Orange
16.	PH of 0.001 M solution of KOH is:	14	13	11	2
17.	Water of crystallization in copper sulphate $CuSO_4$ is:	$5 H_2O$	$2H_2O$	$3H_2O$	$10H_2O$
18.	The water of crystallization is responsible for the:	Melting point of crystals	Shapes of crystals	Boiling points of crystals	Transition point of crystals

(Section - I)

2. Write short answers to the following questions.

(11×2=22)

- i. Give Arrhenius concept of acid and bases.
- ii. Write characteristics of acids.
- iii. Why water is called as an amphoteric specie?
- iv. Why H^+ ion act as Lewis acid?
- v. How bases react with ammonium salt?
- vi. Define an acid and a base according to Bronsted-Lowry concept.
- vii. What are the limitations of Arrhenius concept?
- viii. A solution of hydrochloric acid is 0.01 M. What is its PH value?
- ix. Define PH. What is the PH of pure water?
- x. Define universal indicator.
- xi. Write use of sulphuric acid and nitric acid.

(Section - II)

Note: Give detailed answers of the following questions.

(5+4=09)

3. a) Write a detailed note on Lewis concept of Acid and Bases. (5)
- b) How bases precipitate insoluble hydroxides? (4)