Adabistan-e-Soophia

Test Session (2020 – 21)
Code: 1079

Test No.: **10**   **Paper:** Chemistry

Name:_____   **Class:** IX **Sec:** _____

Syllabus: **Full Book**

Question Numbers	1	2	3	4	5	6	7	Total	Grade	%age
Maximum Marks	12	10	10	10	09	09	09	**60**		
Marks Obtained										

Remarks:_____

_____

Sheet No.

Roll Number Sheet

| | | | | | | | |
|D|D|M|M|Y|Y| | |

Date: ⬚⬚⬚⬚⬚⬚   Morning Ⓜ
                  Evening Ⓔ

Roll No.

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③③③③③③
④④④④④④
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Paper Code

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⑧⑧⑧⑧
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ہدایات

● امیدوار کو صرف نیلے یا کالے پین / مارکر استعمال کرنے کی اجازت ہے۔
● اس بات کا خاص خیال رکھیں کہ دائرہ مکمل پُر ہو اور سیاہی دائرے سے باہر نہ نکلے۔
مثال: (۱) صحیح ● (۲) غلط ✹ (۳) غلط ◖
● کاغذ کو موڑنا یا سٹیپل کرنا منع ہے۔
● دائروں کے اوپر دی گئی مخصوص جگہ پر Roll No اور Paper Code لکھیے اور نیچے دیے گئے دائروں کو اس طرح پُر کریں کہ ہر خانے میں ایک ہندسہ آئے۔
● **نوٹ:**
ایک سے زیادہ دائروں کو پُر کرنے / کاٹ کر پُر کرنے کی صورت میں مذکورہ جواب / رول نمبر / پیپر کوڈ غلط تصور ہو گا جس کی تمام ذمہ داری طالب علم پر ہو گی

Date	
Matric	Ⓜ
Inter	Ⓘ
Part 1	①
Part 2	②
Annual	Ⓐ
Supply	Ⓢ
Morning	Ⓜ
Evening	Ⓔ
Subject	

INSTRUCTIONS

Use BLUE/BLACK ball pen or marker ONLY.

Fill/ Shade the circle completely.

EXAMPLE (I) CORRECT ●   (II) INCORRECT ✹   (III)INCORRECT ◖

Folding, Stapling, Cutting and putting unnecessary sign and identification on the sheet is strictly NOT allowed.

Please write the Roll No, Paper Code in the space allocated above the circles and carefully fill/ shade the circles.

DISCLAIMER: IN CASE OF INCOMPLETE/ CANCELLATION/ MULTIPLE FILLING OF CIRCLES, THE ANSWER SHALL BE DECLARED AS WRONG.

Name:

Father's Name:

Roll Number (in words)

Signature of Candidate:

Signature of Deputy Superintendent:

MCQs RESPONSE PART (TO BE FILLED BY THE STUDENT) (اُمیدوار خود پُر کرے)

	A	B	C	D			A	B	C	D
1.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		13.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
2.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		14.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
3.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		15.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
4.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		16.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
5.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		17.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
6.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		18.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
7.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		19.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
8.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		20.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
9.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		21.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
10.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		22.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
11.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		23.	Ⓐ	Ⓑ	Ⓒ	Ⓓ
12.	Ⓐ	Ⓑ	Ⓒ	Ⓓ		24.	Ⓐ	Ⓑ	Ⓒ	Ⓓ

Paper Code

0 0 0 0
1 1 1 1
2 2 2 2
3 3 3 3
4 4 4 4
5 5 5 5
6 6 6 6
7 7 7 7
8 8 8 8
9 9 9 9

نوٹ:

اـ ایک سے زیادہ دائروں کو پُر کرنے یا کاٹ
کر پُر کرنے کی صورت میں مذکورہ جواب غلط
تصور ہو گا۔

Marks Obtained in MCQs _____ (To be filled by the examiner only) (صرف ممتحن پُر کرے)

AWARD SHEET (FOR OFFICIAL USE IN MARKING CENTRE) (صرف ممتحن پُر کرے)

MCQs Marks	Q 2 Marks	Q 3 Marks	Q 4 Marks	Q 5 Marks	Q 6 Marks	Q 7 Marks	Q 8 Marks	Q 9 Marks	Q 10 Marks	Total Marks
0 0	0 0	0 0	0 0	0 0	0 0	0 0	0 0	0 0	0 0	0 0 0
1 1	1 1	1 1	1 1	1 1	1 1	1 1	1 1	1 1	1 1	1 1 1
2 2	2 2	2 2	2 2	2 2	2 2	2 2	2 2	2 2	2 2	2 2 2
3 3	3 3	3 3	3 3	3 3	3 3	3 3	3 3	3 3	3 3	3 3 3
4 4	4 4	4 4	4 4	4 4	4 4	4 4	4 4	4 4	4 4	4 4 4
5 5	5 5	5 5	5 5	5 5	5 5	5 5	5 5	5 5	5 5	5 5 5
6 6	6 6	6 6	6 6	6 6	6 6	6 6	6 6	6 6	6 6	6 6 6
7 7	7 7	7 7	7 7	7 7	7 7	7 7	7 7	7 7	7 7	7 7 7
8 8	8 8	8 8	8 8	8 8	8 8	8 8	8 8	8 8	8 8	8 8 8
9 9	9 9	9 9	9 9	9 9	9 9	9 9	9 9	9 9	9 9	9 9 9

Sub-Examiner Code

0 0 0 0 0
1 1 1 1 1
2 2 2 2 2
3 3 3 3 3
4 4 4 4 4
5 5 5 5 5
6 6 6 6 6
7 7 7 7 7
8 8 8 8 8
9 9 9 9 9

It is certified that the result on this form matches
with marks given on the corresponding answer sheet.

Sub-Examiner Signature _____

**DISCLAIMER: BOARD SHALL NOT BE RESPONSIBLE IF THE MENTIONED
INSTRUCTIONS ARE NOT FOLLOWED.**

Note: Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more times result in zero mark in that question.

Q.1	Questions	(A)	(B)	(C)	(D)
1.	Number of moles present in 0.4g of H_2SO_4 is:	0.22	0.025	0.00408	0.0408
2.	Molar mass of CO_2 is:	44 amu	12g	40g	44g
3.	Electronic configuration of Cl^- is:	$1s^2 2s^2 2p^6 3s^2 3p^5$	$1s^2 2s^2 2p^6 3s^2 3p^6$	$1s^2 2s^2 2p^6$	$1s^2 2s^2 2p^5$
4.	Pressure inside the discharge tube is kept:	1 atm	10 atm	10^{-4} atm	10^4 atm
5.	Elements in the start of periods are:	more metallic	less metallic	non metallic	more non metallic
6.	In rusting process oxidation occurs at:	cathodic region	anodic region	neutral region	both (A) & (B)
7.	In electroplating of silver, anode is made up of:	silver	copper	antimonial lead	chromium
8.	Process of evaporation increases with the increase in:	pressure	temperature	intermolecular forces	size of molecules
9.	Triple covalent bond is present in:	CH_4	C_2H_2	C_2H_4	Cl_2
10.	The modern periodic table is based on:	electronic number	proton number	neutron number	mass number
11.	An atom can lose its electron if they are:	more than 3	less than 3	1, 2 or 3	both (B) & (C)
12.	More reactive metal is:	sodium	cesium	potassium	lithium

(Section - I)

2. **Write short answers to the following questions. (any five)**   **(5×2=10)**

    i.   Define elements.

    ii.   Differentiate between organic and inorganic chemistry.

    iii. What is an Avogadro number?

    iv.   What are the conditions for writing electronic configuration?

    v.   Why energy of a revolving electron is quantized?

    vi.   Define transition elements.

    vii. What are the trends of electron affinity along the periodic table?

    viii. What is meant by blocks of periodic table?

3. **Write short answers to the following questions. (any five)**   **(5×2=10)**

    i.   Define double covalent bond.

    ii.   Why is BF_3 an electron deficit?

    iii. What are intermolecular forces? Compare these forces with chemical bond forces with reference to HCl molecule?

    iv.   A sample of oxygen gas has a volume of $250cm^3$ at $-30°C$. If a gas is allowed to expand up to $700cm^3$ at constant pressure, find out its final temperature.

    v.   Does process of evaporation occur only at high temperature. Explain with example.

    vi.   How can we distinguish between a solution and a pure liquid?

    vii. Differentiate between solute and solvent.

    viii. How much NaOH is required to prepare its $500cm^3$ of 0.4M solution?

4. **Write short answers to the following questions. (any five)**   **(5×2=10)**

    i.   Define galvanizing.

    ii.   Define oxidation reactions.

    iii. Describe the process of alloying.

    iv.   Calculate the oxidation number of sulfur in H_2SO_4.

    v.   Define the term inertness of noble metals.

    vi.   Define catalytic converter.

    vii. Define nonmetals. Give names of nonmetals of group 17.

    viii. Why is copper used to make electrical wires.

(Section - II)

Note: Give detailed answers of the following questions. (any two)   **(9x2=18)**

5. (a) Define the term valency. Explain valency of covalent and ionic compounds.   **(5)**

(b) Differentiate between Rutherford's and Bohr's atomic model.   **(4)**

6. (a) Write a note on hydrogen bonding.   **(5)**

(b) Define vapor pressure. What are the factors affecting it.   **(4)**

7. (a) Define concentration. What are the various types of units used to express concentration of solutions?   **(5)**

(b) How electroplating of chromium is carried out?   **(4)**